

Name: \_\_\_\_\_

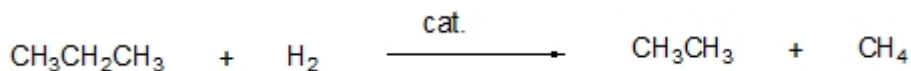
**Organic I Lecture**

**Fall 2012**

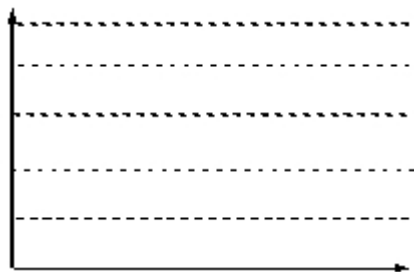
**Quiz #3**

(10 points)

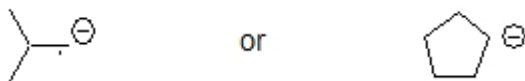
1. Use the table of bond dissociation enthalpies on the back of this page, to calculate the values of  $\Delta H^\circ$  in Kcal/mol for the following reaction. State whether the reaction is endothermic or exothermic. Show all of your work for full credit. (3 points, problem 4-40d)



2. Provide an energy diagram for a two step exogonic reaction where the first step is rate determining and was found to have a late transition state. Correctly label each axis. (4 points)



3. Circle the more stable carbanion. (1 points)



4. Draw a structure for the major monobromination product that would result from treating 2,2-dimethylbutane with bromine and light. (2 points)